

Chapter: 1

Practice sheet

Q1: Translate the following statements into chemical equations and then balance them:

- (a) Barium chloride reacts with aluminium sulphate to give aluminium chloride and a precipitate of barium sulphate.
- (b) Potassium metal reacts with water to give potassium hydroxide and hydrogen gas.
- (c) Solutions of barium chloride and sodium sulphate in water react to give insoluble barium sulphate and the solution of sodium chloride. (3 marks)

Q2: Write the balanced equation for the following chemical reactions

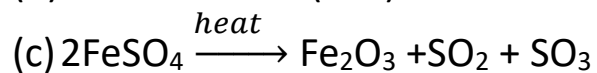
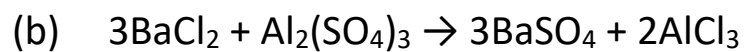
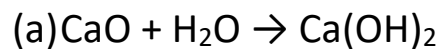
- (a) Barium chloride + Aluminium sulphate \rightarrow Barium sulphate + Aluminium chloride
- (b) Aluminium + Copper chloride \rightarrow Aluminium chloride + Copper
- (c) Calcium hydroxide + Carbon dioxide \rightarrow Calcium carbonate + Water (3 marks)

Q3: Explain the following in terms of gain or loss of oxygen with two examples each. (a) Oxidation (b) Reduction. (3 marks)

Q4: What is the difference between displacement and double displacement reactions? Write equations for these reactions. (2 marks)

Q5: Why are decomposition reactions called the opposite of combination reactions? Write equations for these reactions. (2 marks)

Q6: Name the type of chemical reaction represents by the following equation:



(3 marks)



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