Practice test
Chapter: 1st
Chemical Reactions and Equations

Q1: Which of the statements about the reaction below are incorrect?

$$2PbO(s) + C(s) \rightarrow 2Pb(s) + CO2(g)$$

(a) Lead is getting reduced.
(b) Carbon dioxide is getting oxidised.
(c) Carbon is getting oxidised.
(d) Lead oxide is getting reduced.

(i) (a) and (b)
(ii) (a) and (c)
(iii) (a), (b) and (c)
(iv) all

Q2: What chemicals are used in fireworks?
(A) Copper chloride
(B) Calcium chloride
(C) Barium chloride
(D) All of above

Q3: Consider equations:
$Ca^{+2}(aq) + 2OH^{-}(aq) \rightarrow Ca(OH)_2(s)$. Precipitate of calcium hydroxide will be of
(A) Green color
(B) Blue color
(C) Brown color
(D) White color

Q4: Translate the following statements into chemical equations and then balance them.
(a) Barium chloride reacts with aluminium sulphate to give aluminium chloride and a precipitate of barium sulphate.
(b) Aluminium + Copper chloride → Aluminium chloride + Copper

Q5: Why is respiration considered an exothermic reaction? Explain.
Q6: What do you mean by a precipitation reaction? Explain by giving examples.
Q7: Oil and fat containing food items are flushed with nitrogen. Why?

Q8: **Assertion (A):** When zinc rod is dipped into aqueous solution of copper sulphate, the colour of the solution changes.

Reason (R): Zinc being more reactive displaces less reactive. Copper from its aqueous solution resulting in the formation of zinc sulphate which is colourless.

Q9: **Assertion (A):** Photosynthesis is an endothermic reaction.

Reason (R): Energy from sunlight is absorbed in photosynthesis.

Q10: Why does the colour of copper sulphate solution change when an iron nail is dipped in it?