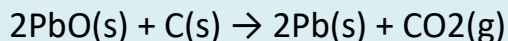


Practice test
Chapter: 1st
Chemical Reactions and Equations

Q1: Which of the statements about the reaction below are incorrect?

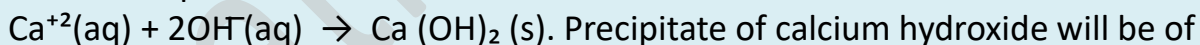


- (a) Lead is getting reduced.
 - (b) Carbon dioxide is getting oxidised.
 - (c) Carbon is getting oxidised.
 - (d) Lead oxide is getting reduced.
- (i) (a) and (b)
 - (ii) (a) and (c)
 - (iii) (a), (b) and (c)
 - (iv) all

Q2: What chemicals are used in fireworks?

- (A) Copper chloride
- (B) Calcium chloride
- (C) Barium chloride
- (D) All of above

Q3: Consider equations:



- (A) Green color
- (B) Blue color
- (C) Brown color
- (D) White color

Q4: Translate the following statements into chemical equations and then balance them.

- (a) Barium chloride reacts with aluminium sulphate to give aluminium chloride and a precipitate of barium sulphate.
- (b) Aluminium + Copper chloride $\rightarrow$ Aluminium chloride + Copper

Q5: Why is respiration considered an exothermic reaction? Explain.

Q6: What do you mean by a precipitation reaction? Explain by giving examples.

Q7: Oil and fat containing food items are flushed with nitrogen. Why?

Q8: Assertion (A): When zinc rod is dipped into aqueous solution of copper sulphate, the colour of the solution changes.

Reason (R): Zinc being more reactive displaces less reactive. Copper from its aqueous solution resulting in the formation of zinc sulphate which is colourless.

Q9: Assertion (A): Photosynthesis is an endothermic reaction.

Reason (R): Energy from sunlight is absorbed in photosynthesis.

Q10: Why does the colour of copper sulphate solution change when an iron nail is dipped in it?

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